



# Bis(diphenylphosphino)acetylene as bifunctional ligand in cycloplatinated complexes: Synthesis, characterization, crystal structures and mechanism of MeI oxidative addition

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## ABSTRACT

Binuclear cycloplatinated(II) complexes with general formula of  $[\text{Pt}_2\text{Me}_2(\text{CN})_2(\mu\text{-dppac})]$ , (**1**, CN = deprotonated 2-phenylpyridine (ppy); **2**, CN = deprotonated benzo[h]quinoline (bhq)) in which dppac = 1,1'-bis(diphenylphosphino)acetylene, are synthesized by the reaction of  $[\text{PtMe}(\text{SMe}_2)(\text{CN})]$  with 0.5 equiv of dppac at room temperature. The complexes are fully characterized using multinuclear ( $^1\text{H}$ ,  $^{31}\text{P}$  and  $^{195}\text{Pt}$ ) NMR spectroscopy and complex **2** is further identified by single crystal X-ray structure determination. Kinetics of the reaction of complexes **1** and **2** with MeI are investigated in  $\text{CHCl}_3$  and based on the UV–vis and  $^{31}\text{P}$  NMR data, a mechanism involving a double MeI oxidative addition is suggested. The classical  $\text{S}_{\text{N}}2$  mechanism is proposed for both steps and the involved intermediates are suggested. Although MeI in each step was *trans* oxidatively added to one of the platinum(II) centers, further *trans* to *cis* isomerizations of methyl and iodide ligands were also identified. The rates are almost four times slower in the second step as compared to the first step due to the electronic effects transmitted through the dppac ligand. Reaction rates concerning complex **2**, having bhq ligand, are almost 1.3 times slower than those involving the related ppy complex **1**. This is attributed to the stronger donor ability of the ppy ligand, as compared to that bhq ligand and is in agreement with the values of  $^1J_{\text{PtP}}$  observed in the  $^{31}\text{P}$  NMR spectra of the complexes **1** and **2**. The structure of the Pt(IV)–Pt(IV) dimeric complex  $[\text{Pt}_2\text{I}_2\text{Me}_4(\text{ppy})_2(\mu\text{-dppac})]$ , **3**, produced by oxidative addition of complex **1** to MeI, is also determined using X-ray diffraction which is the first X-ray structural determination of a diplatinum complex containing two Pt(IV) centers bridged by one dppac ligand.

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## 1. Introduction

Binuclear transition metal complexes that show ligand-mediated metal-to-metal electronic communications are extensively investigated due to their potential application to molecular-scale electronics and devices [1,2]. Additionally, these complexes are attractive because they provide interesting mixed-valence species that are regarded as prototypes for molecular switches and wires [1–9]. The binuclear complexes containing bridging ligands have

also received much attention due to their potential applications in catalysis, molecular recognition and photoactive devices [10–12]. A key to synthesize the related two-centered metal-based material is the selection of bridging ligands capable of mediating electronic communications. The spacers used are, in most cases, either C-donor or N-/O-donor ligands connected by *sp* or *sp*<sup>2</sup> carbon chains [13–20]. Like the conjugated N- or C-donor ligands, the rodlike P-donor spacers also display some characteristics such as photostability, conjugation and rigidity. These P-donor ligands have molecular orbitals with suitable energies to overlap with those of the attached metal centers. The capacity of bis(diphenylphosphino)acetylene ( $\text{Ph}_2\text{PC}\equiv\text{CPh}_2$ , abbreviated as dppac) to act as a bifunctional ligand to transition metals has been examined [21–28]. This ligand has a

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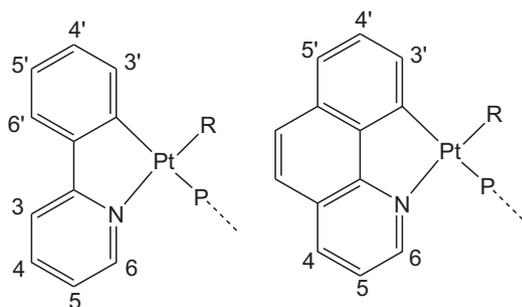
linear geometry with two terminal phosphine moieties and one acetylene moiety able to coordinate to metal ions [29,30]. The dominant coordination modes, which leave the ligand intact, are end-on terminal [31,32] and end-on bridging [33–36] coordination through the phosphorus atoms.

On the other hand, transition metal cyclometalated complexes, in particular those involving platinum and 2-phenylpyridine are of interest due to their potential applications as chemosensors [37], photocatalysts [38] and luminescence [39–50]. Square planar cyclometalated platinum complexes have also been used as “building blocks” for complex systems such as self-assembly [51] and dendrimers [52,53].

We have recently studied the synthesis and reactivity of some cyclometalated organoplatinum complexes aiming to prepare more complex structures for potential applications as molecular materials. Our previous works describe the preparation of some types of complexes with bidentate phosphine ligand, including bis(diphenylphosphino)methane, dppm, bis(diphenylphosphino)ethane, dppe and 1,1'-bis(diphenylphosphino)ferrocene, dppf [54,55]. For the purpose of designing new complexes potentially suitable for molecular wires based on platinum complexes, we have examined to the construction of novel dimers of cyclometalated platinum(II) complexes linked by the conjugated rigid diphosphine spacer, bis(diphenylphosphino)acetylene (dppac). The precursors used for this purpose are [PtMe(CN)(SMe<sub>2</sub>)] (CN = deprotonated 2-phenylpyridine (ppy) or deprotonated benzo{h}quinoline (bhq)). In this paper, we describe the preparation, structural and spectroscopic characterization of binuclear cyclometalated platinum complexes connected by conjugated diphosphine, dppac. We also describe here the kinetic and mechanism of the reaction of Pt(II)–Pt(II) complexes with MeI to form new Pt(IV)–Pt(IV) complexes, the first binuclear cyclometalated Pt(IV) complexes containing dppac as bridging ligand.

## 2. Experimental section

The <sup>1</sup>H, <sup>13</sup>C, <sup>31</sup>P and <sup>195</sup>Pt NMR spectra were recorded on a Bruker Avance DRX 500 MHz spectrometer in CDCl<sub>3</sub> as solvent. The operating frequencies and references, respectively, are shown in parentheses as follows: <sup>1</sup>H (500 MHz, TMS), <sup>13</sup>C (125 MHz, TMS), <sup>31</sup>P (202 MHz, 85% H<sub>3</sub>PO<sub>4</sub>), and <sup>195</sup>Pt (107 MHz, aqueous Na<sub>2</sub>PtCl<sub>4</sub>). The chemical shifts and coupling constants are in ppm and Hz, respectively. Kinetic studies were carried out by using a Perkin–Elmer Lambda 25 spectrophotometer with temperature control using an EYELA NCB-3100 constant temperature bath. Benzo{h}quinoline, 2-phenylpyridine and bis(diphenylphosphino)acetylene were commercially available and the precursor compounds [PtMe(CN)(SMe<sub>2</sub>)] (CN = deprotonated 2-phenylpyridine (ppy) or deprotonated benzo{h}quinoline (bhq)) were prepared according to reported procedures [54,56,57]. Scheme 1 shows the structural



**Scheme 1.** Representation of CN ligands [ppy (left) and bhq (right)] with position labeling.

formula of the CN ligands with the carbon atoms numbered to clarify the following chemical shift assignments.

### 2.1. [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **1**

To a solution of [PtMe(ppy)(SMe<sub>2</sub>)] (100 mg, 0.23 mmol) in acetone (15 mL) was added 0.5 equiv bis(diphenylphosphino)acetylene (46 mg, 0.12 mmol) at room temperature. The mixture was stirred for 3 h. The solvent was removed under reduced pressure, and the residue was washed with diethyl ether (2 × 3 mL) and dried under vacuum. Yield 100 mg; 76%, mp. 230 °C (decomp). Anal. Calcd. for C<sub>50</sub>H<sub>42</sub>N<sub>2</sub>P<sub>2</sub>Pt<sub>2</sub>: C, 53.5; H, 3.8; N, 2.5; Found: C, 53.8; H, 3.6; N, 2.3. NMR data: δ(<sup>1</sup>H) 0.70 [d, <sup>3</sup>J<sub>PH</sub> = 8.9 Hz, <sup>2</sup>J<sub>PtH</sub> = 82.5 Hz, 6H, Me ligands], 6.40 [ddd, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> = 5.8 Hz, <sup>3</sup>J<sub>H<sup>4'</sup>H<sup>5'</sup></sub> = 8.6 Hz, <sup>5</sup>J<sub>PH</sub> = 1.3 Hz, 2H, H<sup>4'</sup> of ppy], 7.10–7.37 (overlapping hydrogens), 7.42 [t, <sup>3</sup>J<sub>H<sup>5'</sup>H<sup>6'</sup></sub> ≈ <sup>3</sup>J<sub>H<sup>4'</sup>H<sup>5'</sup></sub> = 8.2 Hz, 2H, H<sup>5'</sup> of ppy], 7.81 [td, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> ≈ <sup>4</sup>J<sub>PH</sub> ≈ 6.0 Hz, <sup>4</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> = 1.4 Hz, <sup>3</sup>J<sub>PtH</sub> = 54.1 Hz, 2H, H<sup>3'</sup> of ppy], 8.30 [d, <sup>3</sup>J<sub>H<sup>5'</sup>H<sup>6'</sup></sub> = 5.9 Hz, <sup>3</sup>J<sub>PtH</sub> = 14.8 Hz, 2H, H<sup>6'</sup> of ppy], δ(<sup>31</sup>P) 12.6 [s, <sup>1</sup>J<sub>PtP</sub> = 2049 Hz, 2P of dppac]; δ(<sup>195</sup>Pt) –2594 [d, <sup>1</sup>J<sub>PtP</sub> = 2047 Hz, 2Pt].

### 2.2. [Pt<sub>2</sub>Me<sub>2</sub>(bhq)<sub>2</sub>(μ-dppac)], **2**

This was prepared by the method as described above for preparation of complex **1** using the starting materials [PtMe(bhq)(SMe<sub>2</sub>)] and dppac. Yield: 74%, mp. 250 °C (decomp). Anal. Calcd. for C<sub>54</sub>H<sub>42</sub>N<sub>2</sub>P<sub>2</sub>Pt<sub>2</sub>: C, 55.4; H, 3.6; N, 2.4; Found: C, 55.6; H, 3.5; N, 2.2. NMR data: δ(<sup>1</sup>H) 0.88 [d, <sup>2</sup>J<sub>PtH</sub> = 82.7 Hz, <sup>3</sup>J<sub>PH</sub> = 9.0 Hz, 6H, Me ligands], 6.60 [dd, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> = 6.1 Hz, <sup>3</sup>J<sub>H<sup>4'</sup>H<sup>5'</sup></sub> = 8.0 Hz, 2H, H<sup>4'</sup> of bhq], 7.25–7.76 (overlapping hydrogens), 8.10 [td, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> ≈ <sup>4</sup>J<sub>PH</sub> ≈ 6.1 Hz, <sup>3</sup>J<sub>H<sup>5'</sup>H<sup>6'</sup></sub> = 1.4 Hz, <sup>3</sup>J<sub>PtH</sub> = 51.8 Hz, 2H, H<sup>3'</sup> of bhq], 8.50 [dt, <sup>3</sup>J<sub>H<sup>5'</sup>H<sup>6'</sup></sub> = 5.3 Hz, <sup>4</sup>J<sub>HP</sub> ≈ <sup>4</sup>J<sub>H<sup>4'</sup>H<sup>5'</sup></sub> ≈ 1.4 Hz, <sup>3</sup>J<sub>PtH</sub> = 14.7 Hz, 2H, H<sup>6'</sup> of bhq]; δ(<sup>31</sup>P) 12.3 [s, <sup>1</sup>J<sub>PtP</sub> = 2102 Hz, 2P of dppac]; δ(<sup>195</sup>Pt) –2590 [d, <sup>1</sup>J<sub>PtP</sub> = 2104 Hz, 2Pt].

### 2.3. [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **3**

To a solution of [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **1**, (50 mg in 25 mL CH<sub>2</sub>Cl<sub>2</sub>) was added an excess of MeI (550 μL) at room temperature and the mixture was stirred for 2 h. The solvent was removed under reduced pressure, and the residue was washed with diethyl ether (2 × 3 mL). The product was dried under vacuum. Yield 40 mg; 64%, mp. 175 °C (decomp.). Anal. Calcd. for C<sub>52</sub>H<sub>48</sub>I<sub>2</sub>N<sub>2</sub>P<sub>2</sub>Pt<sub>2</sub>: C, 44.4; H, 3.4; N, 2.0; Found: C, 44.2; H, 3.2; N, 1.7. NMR data: δ(<sup>1</sup>H) 1.16 (d, <sup>3</sup>J<sub>PH</sub> = 7.9 Hz, <sup>2</sup>J<sub>PtH</sub> = 62.1 Hz, 3H, 1 Me group *trans* to P), 1.17 (d, <sup>3</sup>J<sub>PH</sub> = 7.9 Hz, <sup>2</sup>J<sub>PtH</sub> = 62.0 Hz, 3H, 1 Me groups *trans* to P), 1.50 (d, <sup>3</sup>J<sub>PH</sub> = 8.7 Hz, <sup>2</sup>J<sub>PtH</sub> = 70.1 Hz, 3H, 1 Me group *trans* to N), 1.52 (d, <sup>3</sup>J<sub>PH</sub> = 8.9 Hz, <sup>2</sup>J<sub>PtH</sub> = 70.3 Hz, 3H, 1 Me group *trans* to N), (aromatic protons): 6.8–7.9 (overlapping multiplets), 9.42 (d, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> = 5.2 Hz, <sup>3</sup>J<sub>PtH</sub> = 9 Hz, 1H, CH group adjacent to coordinated N atom), 9.47 (d, <sup>3</sup>J<sub>H<sup>3'</sup>H<sup>4'</sup></sub> = 5.2 Hz, <sup>3</sup>J<sub>PtH</sub> = 9 Hz, 1H, CH group adjacent to coordinated N atom); δ(<sup>13</sup>C) –5.9 (d, <sup>1</sup>J<sub>PtC</sub> = 628 Hz, <sup>2</sup>J<sub>PC</sub> = 23.6 Hz, 1C, Me *trans* to N), –5.7 (d, <sup>1</sup>J<sub>PtC</sub> = 628 Hz, <sup>2</sup>J<sub>PC</sub> = 23.7 Hz, 1C, Me *trans* to N), 6.9 (d, <sup>1</sup>J<sub>PtC</sub> = 509 Hz, <sup>2</sup>J<sub>PC</sub> = 116 Hz, 1C, Me *trans* to P), 7.1 (d, <sup>1</sup>J<sub>PtC</sub> = 509 Hz, <sup>2</sup>J<sub>PC</sub> = 116 Hz, 1C, Me *trans* to P), 101.7 (d, with <sup>1</sup>J<sub>PC</sub> = 66.2 Hz, 1C, –C≡C), 101.9 (d, with <sup>1</sup>J<sub>PC</sub> = 66.2 Hz, 1C, –C≡C), 147.0, 147.1 (each a doublet, with <sup>2</sup>J<sub>PC</sub> = 13.6 Hz, <sup>1</sup>J<sub>PtC</sub> = 888 Hz, 2C, C atoms of the ppy ligands connected to Pt atom), other aromatic Cs were not assigned; δ(<sup>31</sup>P) –23.9 (s, <sup>1</sup>J<sub>PtP</sub> = 976 Hz, 1P), –23.7 (s, <sup>1</sup>J<sub>PtP</sub> = 977 Hz, 1P); δ(<sup>195</sup>Pt) –1797 and –1806 (2 overlapping d, each with <sup>1</sup>J<sub>PtP</sub> = 975 Hz, 2 Pt).

### 2.4. [Pt<sub>2</sub>Me<sub>2</sub>(bhq)<sub>2</sub>(μ-dppac)], **4**

This was prepared by the method as described above for preparation of complex **3** using the starting material [Pt<sub>2</sub>Me<sub>2</sub>(bhq)<sub>2</sub>(μ-dppac)], **2**. Yield: 64%, mp. 190 °C (decomp.). Anal. Calcd. for

**Table 1**  
Crystal data, data collection, and structure refinement details for complexes **2** and **3**.

	Complex <b>2</b>	Complex <b>3</b>
Formula	C <sub>54</sub> H <sub>42</sub> N <sub>2</sub> P <sub>2</sub> Pt <sub>2</sub>	C <sub>52</sub> H <sub>48</sub> I <sub>2</sub> N <sub>2</sub> P <sub>2</sub> Pt <sub>2</sub>
Formula weight	1171.02	1406.84
Temperature/K	291(2)	100(2)
Crystal system	Triclinic	Monoclinic
Space group	P-1	P2(1)/c
a/Å	9.6765(10)	14.7189(4)
b/Å	9.8839(10)	16.7379(5)
c/Å	23.624(3)	19.4066(6)
vol/Å <sup>3</sup>	2235.8(4)	4725.2(2)
Z	2	4
D(calc)/Mg m <sup>-3</sup>	1.739	1.978
μ/mm <sup>-1</sup>	6.361	7.326
T <sub>min</sub> /T <sub>max</sub>	0.884/1.000	0.234/0.322
R(000)	1132	2664
No. of reflns	15,489	43,031
No. of independent reflns	8049 [R(int) = 0.0542]	9667 [R(int) = 0.0603]
No. of observed reflns	5143	7771
[I > 2σ(I)]		
Goof (F <sup>2</sup> )	0.75	1.122
R1, wR2 [I > 2σ(I)]	0.0285, 0.0475	0.0373, 0.0592
R1, wR2 (all data)	0.0589, 0.0509	0.0556, 0.0631

C<sub>56</sub>H<sub>48</sub>I<sub>2</sub>N<sub>2</sub>P<sub>2</sub>Pt<sub>2</sub>: C, 46.2; H, 3.3; N, 1.9; Found: C, 46.5; H, 3.2; N, 1.7. NMR data: δ(<sup>1</sup>H) 1.14 (d, <sup>2</sup>J<sub>PtH</sub> = 62.2 Hz, <sup>3</sup>J<sub>PH</sub> = 7.9 Hz, 3H, 1 Me group *trans* to P), 1.16 (d, <sup>3</sup>J<sub>PH</sub> = 7.9 Hz, <sup>2</sup>J<sub>PtH</sub> = 62.6 Hz, 3H, 1 Me group *trans* to P), 1.65 (d, <sup>3</sup>J<sub>PH</sub> = 8.7 Hz, <sup>2</sup>J<sub>PtH</sub> = 70.1 Hz, 3H, 1 Me group *trans* to N), 1.67 (d, <sup>3</sup>J<sub>PH</sub> = 8.7 Hz, <sup>2</sup>J<sub>PtH</sub> = 70.0 Hz, 3H, 1 Me group *trans* to N), (aromatic protons): 6.9–8.1 (overlapping multiplets), 9.50 (d, <sup>3</sup>J<sub>HH</sub> = 5.1 Hz, <sup>3</sup>J<sub>PtH</sub> = 11 Hz, 1H, CH group adjacent to coordinated N atom), 9.60 (d, <sup>3</sup>J<sub>HH</sub> = 5.3 Hz, <sup>3</sup>J<sub>PtH</sub> = 11 Hz, 1H, CH group adjacent to coordinated N atom); δ(<sup>13</sup>C) -6.7 (d, <sup>1</sup>J<sub>PtC</sub> = 622 Hz, <sup>2</sup>J<sub>PC</sub> ≈ 3 Hz, 1C, Me *trans* to N), -6.5 (d, <sup>1</sup>J<sub>PtC</sub> = 624 Hz, <sup>2</sup>J<sub>PC</sub> = 3.7 Hz, 1C, Me *trans* to N), 6.2 (d, <sup>1</sup>J<sub>PtC</sub> = 505 Hz, <sup>2</sup>J<sub>PC</sub> = 115 Hz, 1C, Me *trans* to P), 6.5 (d, <sup>1</sup>J<sub>PtC</sub> = 504 Hz, <sup>2</sup>J<sub>PC</sub> = 115 Hz, 1C, Me *trans* to P), 101.3 (d, with <sup>1</sup>J<sub>PC</sub> = 66 Hz, 1C, -C≡C), 101.8 (d, with <sup>1</sup>J<sub>PC</sub> = 66 Hz, 1C, -C≡C), other aromatic C's were not assigned; δ(<sup>31</sup>P) -23.9 (s, <sup>1</sup>J<sub>PtP</sub> = 975 Hz, 1P), -23.6 (s, <sup>1</sup>J<sub>PtP</sub> = 976 Hz, 1P); δ(<sup>195</sup>Pt) -1832 and -1823 (2 overlapping d, each with <sup>1</sup>J<sub>PtP</sub> = 978 Hz, 2Pt).

## 2.5. Reaction of [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **1**, with MeI

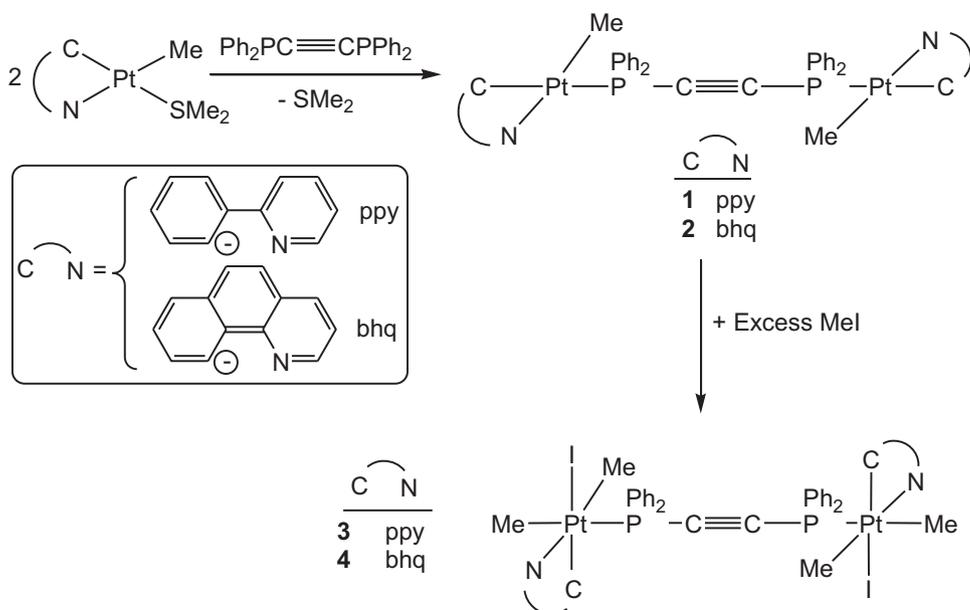
This reaction was monitored by <sup>31</sup>P NMR spectroscopy at room temperature in an NMR tube. To a solution of [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **1**, (10 mg) in CDCl<sub>3</sub> (0.7 mL), an excess of MeI (3 μL) was added. The NMR spectra were recorded several times during about 2 h until the mixture was gradually converted to complex **3** in solution.

## 2.6. Kinetic study

In a typical experiment, a solution of complex **1** in CHCl<sub>3</sub> (3 mL, 1.06 × 10<sup>-4</sup> M) in a cuvette was thermostated at 25 °C and a known excess of MeI was added using a micro syringe. After rapid stirring, the absorbance at λ = 352 nm was monitored with time.

## 2.7. Crystallography

Single crystal of [Pt<sub>2</sub>Me<sub>2</sub>(bhq)<sub>2</sub>(μ-dppac)], **2**, was grown from a concentrated CH<sub>2</sub>Cl<sub>2</sub> solution by slow diffusion of *n*-pentane. Single crystal X-ray data for complex **2** were collected at 296(1) K on STOE IPDS 2T diffractometer (Mo Kα = 0.71073 Å). The cell parameters were retrieved using X-AREA [58] software and refined using X-AREA on all observed reflections. Data reduction and correction for Lp (Lorentz-polarization) and decay were performed using X-AREA software. Absorption corrections were applied using MULABS [59,60]. All structures were solved by direct methods and refined by full-matrix least squares on F<sup>2</sup> for all data using SHELXTL software [61]. All calculations were performed by PLATON. All non-hydrogen atoms were refined anisotropically. Hydrogen atoms were positioned geometrically and refined with a riding model approximation with their parameters constrained to the parent atom with U<sub>iso</sub>(H) = 1.2 or 1.5 U<sub>eq</sub>(C). Single crystal of [Pt<sub>2</sub>I<sub>2</sub>Me<sub>4</sub>(ppy)<sub>2</sub>(μ-dppac)], **3**, was grown from a concentrated CH<sub>2</sub>Cl<sub>2</sub> solution by slow diffusion of *n*-pentane. The single crystal X-ray diffraction study of **3** was carried out on a Bruker Platform D8 CCD diffractometer equipped with Mo Kα radiation (λ = 0.71073). A 0.28 × 0.24 × 0.20 mm



**Scheme 2.** Reactions studied in this work.

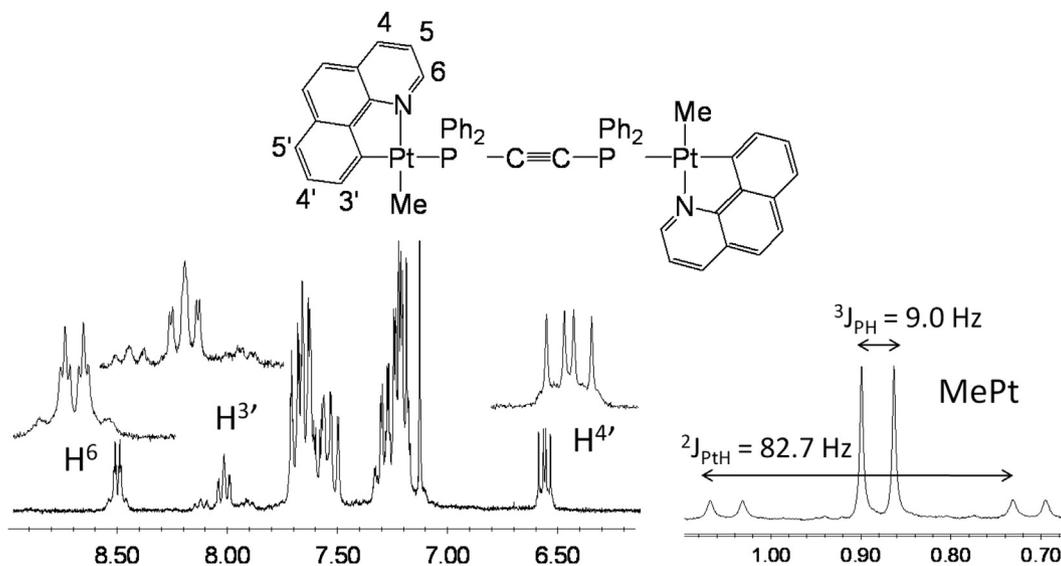


Fig. 1. The  $^1\text{H}$  NMR spectrum of complex **2**.

crystal was mounted on a Cryo-loop with Paratone-N. The Bruker program packages, APEX-II and SAINT were used for data collection, integration and corrections. Solution and refinement proceeded as described for complex **2**. At both Pt(IV) sites of complex **3**, there was iodide/methyl group disorder; the minority component is designated with a prime. The major configuration at Pt1 is 71% occupied, and at Pt2, 88% occupied. The occupancies were refined with a unit occupancy constraint. Crystal data, data collection, and structure refinement details are listed in Table 1.

### 3. Results and discussion

#### 3.1. Synthesis and characterization of the Pt(II)–Pt(II) complexes

The synthetic procedures are summarized in Scheme 2.

The starting complexes  $[\text{PtMe}(\text{CN})(\text{SMe}_2)]$ , in which  $\text{CN} = \text{ppy}$  or  $\text{bhq}$ , reacted with bis(diphenylphosphino)acetylene (dppac) in a 1:0.5 molar ratio to give the organodiplatinum(II) complexes  $[\text{PtMe}_2(\text{CN})_2(\mu\text{-dppac})]$ , **1–2**. They were isolated as stable solids

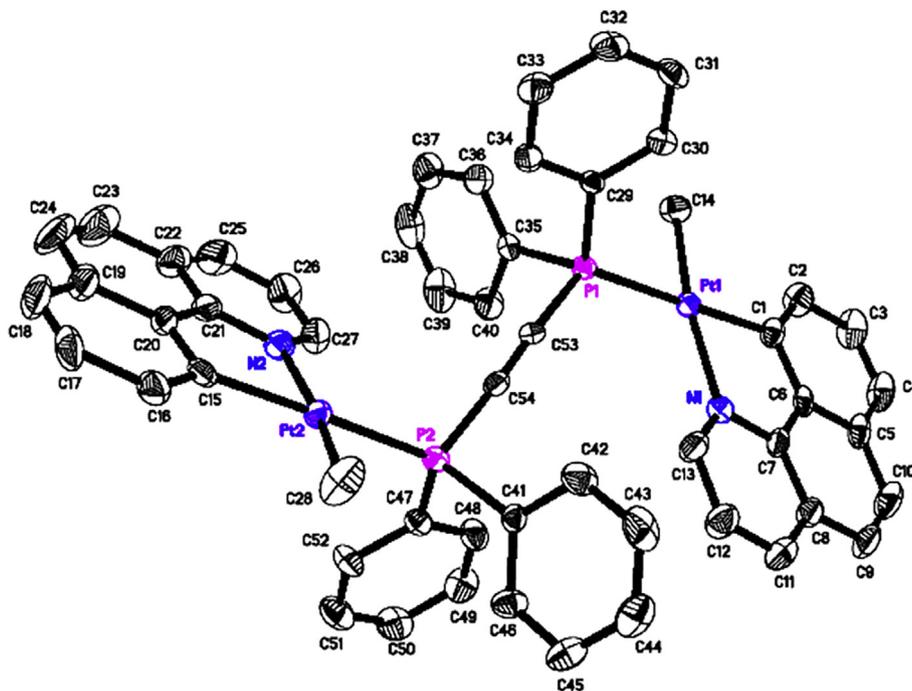


Fig. 2. (a) The molecular structure of complex  $[\text{Pt}_2\text{Me}_2(\text{bhq})_2(\mu\text{-dppac})]$ , **2**, showing 40% probability ellipsoids. The H atoms omitted for clarity. Selected bond distances ( $\text{\AA}$ ) and angles ( $^\circ$ ): Pt(1)–C(1) 2.059(7); Pt(1)–C(14) 2.069(6); Pt(1)–N(1) 2.130(5); Pt(1)–P(1) 2.2672(17); Pt(2)–C(15) 2.053(7); Pt(2)–C(28) 2.066(6); Pt(2)–N(2) 2.144(5); Pt(2)–P(2) 2.2750(18); P(2)–C(54) 1.783(7); P(2)–C(47) 1.824(7); P(2)–C(41) 1.840(6); P(1)–C(53) 1.779(7); P(1)–C(29) 1.836(6); P(1)–C(35) 1.839(6); C(1)–Pt(1)–C(14) 89.6(3); C(1)–Pt(1)–N(1) 80.3(2); C(14)–Pt(1)–N(1) 169.7(2); C(1)–Pt(1)–P(1) 173.85(19); C(14)–Pt(1)–P(1) 85.81(19); N(1)–Pt(1)–P(1) 104.44(15); C(15)–Pt(2)–C(28) 91.2(3); C(15)–Pt(2)–N(2) 80.6(3); C(28)–Pt(2)–N(2) 171.8(3); C(15)–Pt(2)–P(2) 176.7(2); C(28)–Pt(2)–P(2) 91.2(2).

which were characterized by their  $^1\text{H}$ ,  $^{31}\text{P}$  and  $^{195}\text{Pt}$  NMR spectra and elementary analysis and the structure of complex **2** was further determined by X-ray crystallography. Typically, in the  $^1\text{H}$  NMR spectrum of  $[\text{Pt}_2\text{Me}_2(\text{bhq})_2(\mu\text{-dppac})]$ , **2**, in  $\text{CDCl}_3$  at room temperature (see Fig. 1), the methylplatinum resonance appeared at  $\delta = 0.88$  as a doublet due to coupling to phosphorus (with  $^3J_{\text{PH}} = 9.0$  Hz) which is further coupled to  $^{195}\text{Pt}$  (with  $^2J_{\text{PtH}} = 82.7$  Hz) confirming that the methyl ligands are located *trans* to aromatic N atoms, which is characteristic values for Pt–Me *trans* to N [62–64]. The hydrogen related to the CH group adjacent to ligating N atom of the CN ligand appeared as a doublet of triplet at  $\delta = 8.50$  with  $^3J_{\text{HH}} = 5.3$  Hz,  $^4J_{\text{HP}} \approx ^4J_{\text{HN}} \approx 1.4$  Hz and  $^3J_{\text{HPt}} = 14.7$  Hz. The hydrogen related to the CH group adjacent to ligating C atom of the CN ligand appeared as a triplet of doublet at  $\delta = 8.10$  with  $^3J_{\text{HH}} \approx ^4J_{\text{HP}} \approx 6.1$  Hz,

$^4J_{\text{HH}} = 1.4$  Hz and  $^3J_{\text{HPt}} = 51.8$  Hz. In the room temperature  $^{31}\text{P}$  NMR spectrum of the binuclear complex **2**, the observation of a sharp singlet signal for each of the analogs at a chemical shift of 12.3 ppm, which is accompanied by Pt satellites,  $^1J_{\text{PtP}} = 2102$  Hz, confirm that the two PtMe(bhq) moieties, joined together by the dppac spacer ligand, are equivalent. Consistent with this, in the  $^{195}\text{Pt}$  NMR spectrum of **2**, a doublet at  $\delta -2590$  with  $^1J_{\text{PtP}} = 2104$  Hz was observed. This  $^1J_{\text{PtP}}$  value is consistent with P *trans* to a ligand with high trans influence like carbon; *trans* to N would give a larger coupling constant [64–66]. These data suggest that dppac is acting as a spacer ligand between the two PtMe(bhq) moieties and each P atom is coordinated to a Pt atom in a *trans* disposition to coordinating C atom of the phenyl ring of bhq ligand. Each Me ligand is thus located *trans* to the coordinated N atom of bhq ligand.

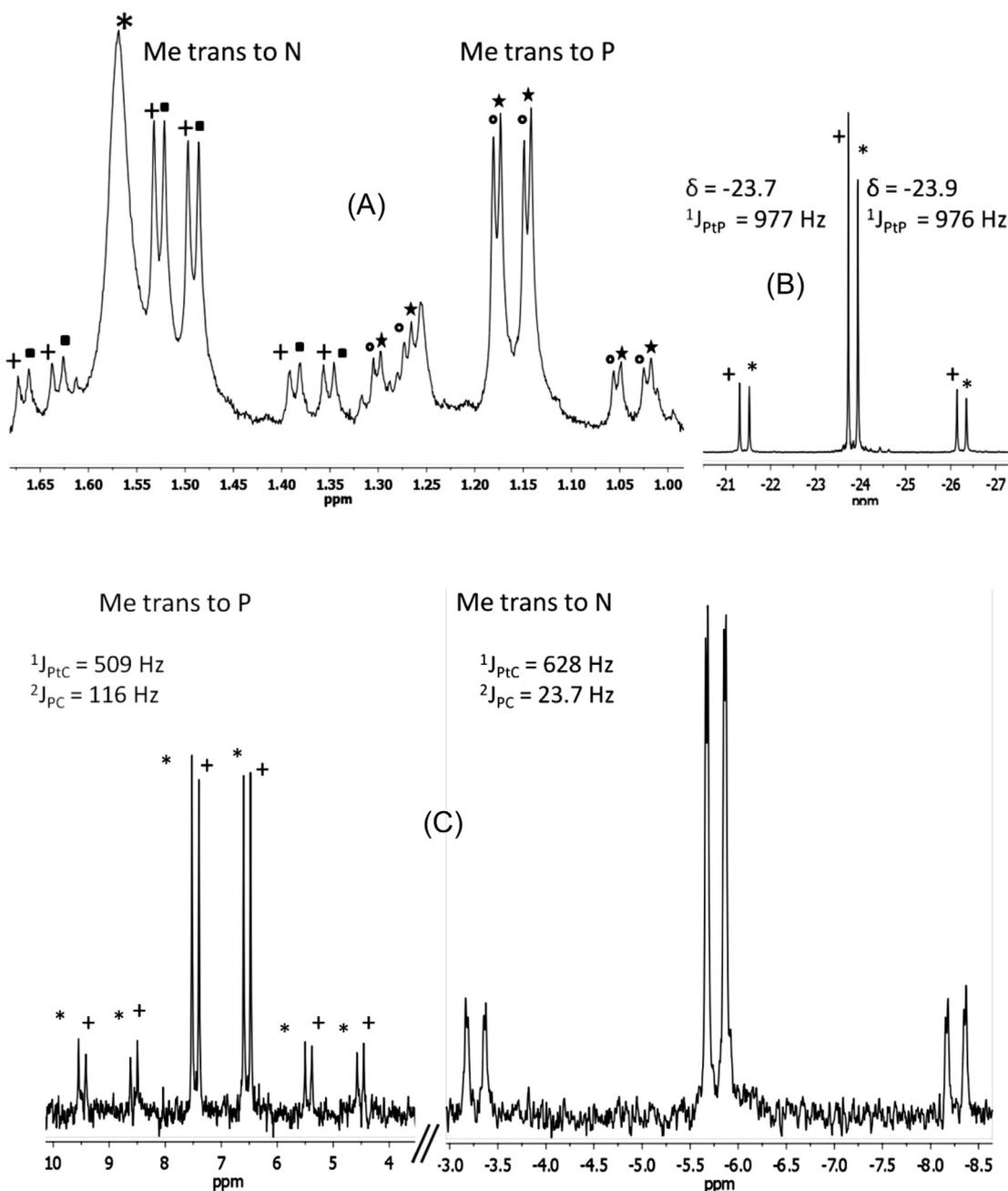
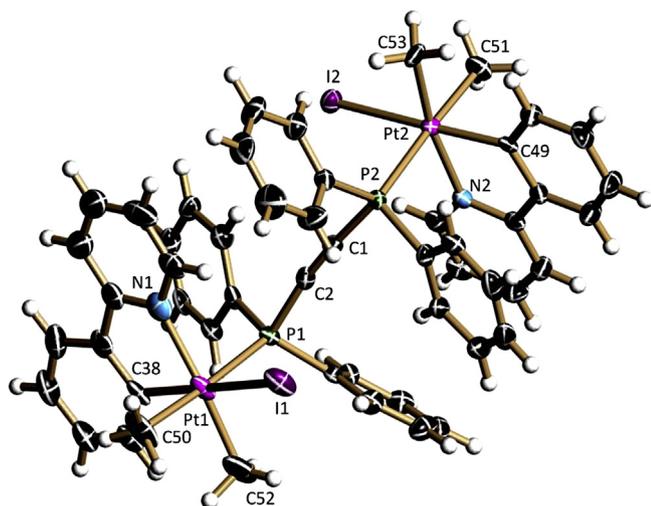


Fig. 3. The NMR spectra of complex **3**. (A):  $^1\text{H}$  in the Me region; (B):  $^{31}\text{P}$  NMR and (C)  $^{13}\text{C}$  NMR in the Me region. The peak labeled \* in  $^1\text{H}$  NMR is due to water of  $\text{CDCl}_3$  solvent.



**Fig. 4.** (a) The molecular structures of complex  $[\text{Pt}_2\text{I}_2\text{Me}_4(\text{ppy})_2(\mu\text{-dppac})]$ , **3**. Selected bond distances (Å) and angles ( $^\circ$ ): Pt1–C38 2.074(6); Pt1–C50 2.102(6); Pt1–C52 2.089(14); Pt1–N1 2.112(6); Pt1–I1 2.7343(7); Pt1–P1 2.3886(14); Pt2–C49 2.037(5); Pt2–C51 2.102(5); Pt2–C53 2.051(9); Pt2–N2 2.133(5); Pt2–I2 2.7378(5); Pt2–P2 2.4146(14); C38–Pt1–C52 92.4(4); C50–Pt1–C52 92.0(4); C38–Pt1–N1 80.3(2); C50–Pt1–N1 89.2(3); C52–Pt1–N1 172.5(4); C38–Pt1–P1 92.41(16); C50–Pt1–P1 176.7(2); N1–Pt1–P1 87.57(13); C38–Pt1–I1 175.12(17); N1–Pt1–I1 97.47(15); P1–Pt1–I1 91.82(4); C49–Pt2–C51 87.5(2); C49–Pt2–N2 79.8(2); C53–Pt2–N2 170.3(3); C51–Pt2–N2 88.7(2); C49–Pt2–P2 92.63(15); C51–Pt2–P2 175.36(18); N2–Pt2–P2 86.80(12); C49–Pt2–I2 175.90(15); C53–Pt2–I2 90.6(2); N2–Pt2–I2 98.08(13); P2–Pt2–I2 90.72(4).

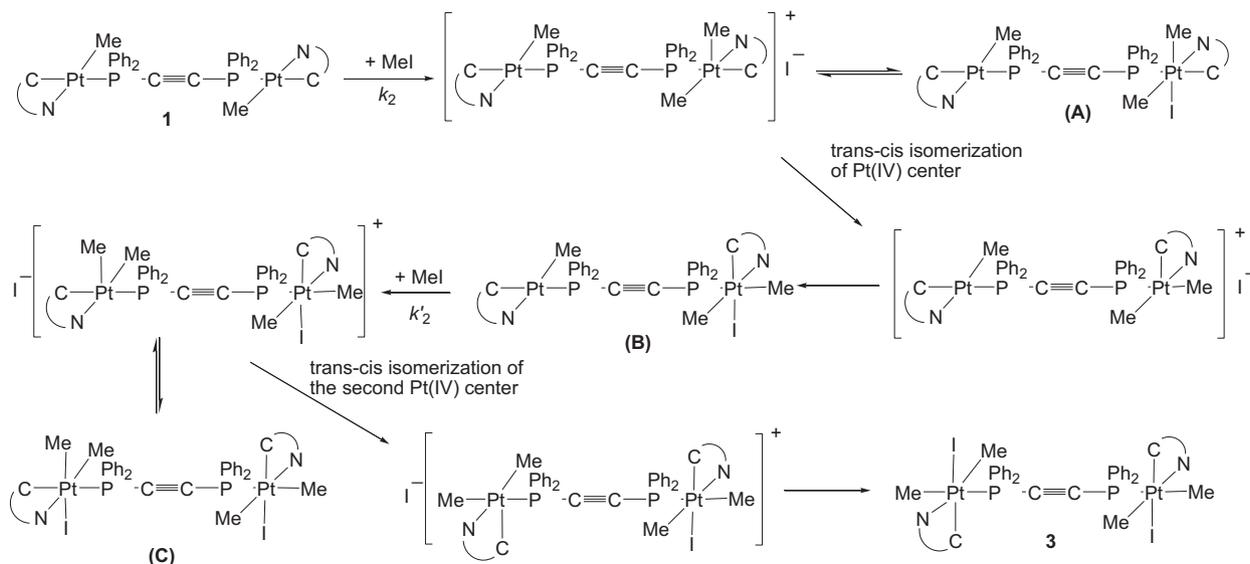
When more than 0.5 equiv of dppac (e.g.,  $[\text{PtMe}(\text{CN})(\text{SMe}_2)]_2:\text{dppac} = 1:1$ ) is used in the related reaction, on the basis of the  $^1\text{H}$  and  $^{31}\text{P}$  NMR spectra, only the corresponding dimer was formed as a mixture with free dppac.

The structure of complex  $[\text{Pt}_2\text{Me}_2(\text{bhq})_2(\mu\text{-dppac})]$ , **2**, was further determined by X-ray crystallography and is shown in Fig. 2; the selected bond lengths and bond angles are also shown. The geometry around the Pt(II) atoms in complex **2** is slightly-distorted square-planar with mean deviation of 0.026(3) Å for Pt1 [C1/N1/C14/P1 plane] and  $-0.021(3)$  Å for Pt2 [C15/N2/C28/P2 plane]. The complex **2** is formed by two identical moieties “PtMe(bhq)”

connected by P atoms of dppac ligand. The fragment P–C–C–P is almost linear [P1–C53–C54 174.4(5) $^\circ$ ] and the Pt1–P1, P1–C53 and Pt1–N1 distances are 2.2672(17), 1.779(7) and 2.130(5) Å, respectively. The terminal Me ligands are in a *transoidal* disposition; the unique C≡C distance is 1.199(7) Å. The square-planar environment for both platinum centers is completed by the aromatic N atom and the *ortho* C of the bhq ligand forming cycloplatinated complex. Mono  $\mu\text{-dppac}$  complexes which a dppac molecule acts as a unique bridging ligand between two metal fragments are not very common [31].

### 3.2. Synthesis and characterization of Pt(IV)–Pt(IV) complexes

The methylplatinum(II) complexes, **1** and **2**, were reacted cleanly with an excess of MeI in  $\text{CH}_2\text{Cl}_2$  at room temperature to give air-stable solid products having the general formula  $[\text{Pt}_2\text{I}_2\text{Me}_4(\text{CN})_2(\mu\text{-dppac})]$  (CN = ppy, **3**; bhq, **4**). The synthesized Pt(IV) complexes were fully characterized using  $^1\text{H}$ ,  $^{13}\text{C}$ ,  $^{31}\text{P}$  and  $^{195}\text{Pt}$  NMR spectroscopy and elemental analysis. Complex **3** was further identified by single crystal X-ray structure determination. Full data are collected in the Experimental section. On the base of  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of complex  $[\text{Pt}_2\text{I}_2\text{Me}_4(\text{ppy})_2(\mu\text{-dppac})]$ , **3**, (see Fig. 3), two different Me groups being *trans* to P and two different Me groups locating *trans* to N ligating atoms were assigned. Thus, in the  $^1\text{H}$  NMR spectrum of complex **3** (see Fig. 3A), the Me groups *trans* to P were observed at  $\delta$  1.16 and 1.17 as doublets with,  $^2J_{\text{PtH}} = 62.1$  and 62.0 Hz, respectively, while Me groups locating *trans* to N ligating atoms were appeared at  $\delta$  1.50 and 1.52, with a considerably higher  $^2J_{\text{PtH}}$  values of 70.1 and 70.3, respectively, due to lower *trans* influence of N atom as compared with that of P atom [66,67]. The hydrogen related to the CH groups adjacent to ligating N atom of the ppy ligands appeared as a doublet at  $\delta = 9.40$  and 9.50 each with  $^3J_{\text{HH}} = 5.2$  Hz and  $^3J_{\text{PtH}} = 9.0$  Hz. Consistently, in the  $^{13}\text{C}$  NMR spectrum of **3**, in the Me region (see Fig. 3C), two doublets were appeared at  $\delta$  6.9 and 7.1 with a  $^2J_{\text{PC}}$  value close to 116 Hz for the Me groups *trans* to P; each doublet further coupled to platinum to give a  $^1J_{\text{PtC}}$  value close to 509 Hz. Two other doublets were appeared further high field at  $\delta -5.9$  and  $-5.7$  each with a  $^2J_{\text{PC}}$  value of close to 23.6 Hz, indicating that they are located *cis* to P atom and thus must be *trans* to N atom; the  $^1J_{\text{PtC}}$  value of 628 Hz for each of the signals is also significantly higher than the value of 508 Hz found

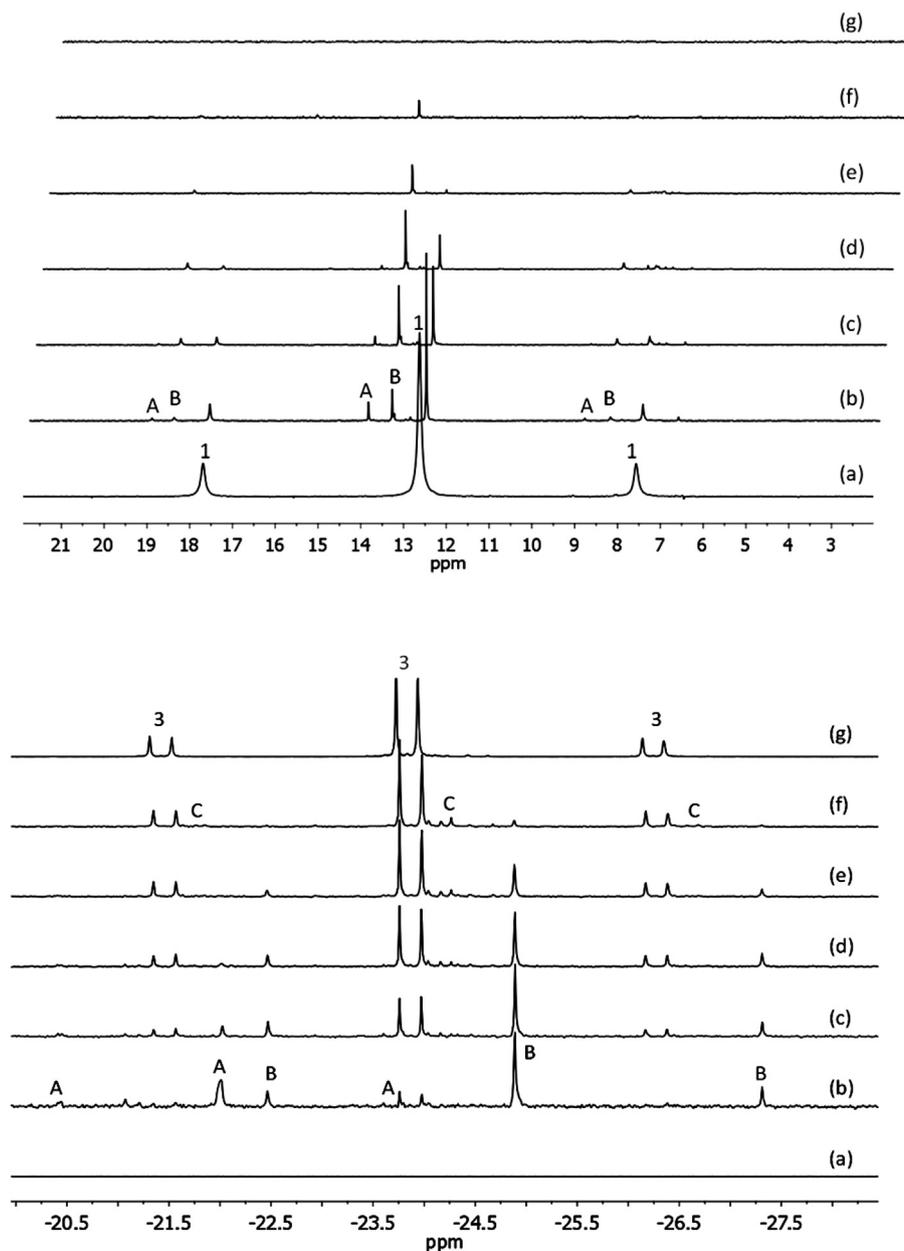


**Scheme 3.** Suggested mechanism for reaction of complexes **1** and **2** with MeI.

for Me groups *trans* to P atoms. In the  $^{31}\text{P}$  NMR spectrum of  $[\text{Pt}_2\text{I}_2\text{Me}_4(\text{ppy})_2(\mu\text{-dppac})]$ , **3**, (see Fig. 3B) two different P signals at  $\delta -23.7$  and  $-23.9$  with  $^1J_{\text{PtP}} = 977$  and  $976$  Hz, respectively, which as expected is much lower than the corresponding value of  $2049$  Hz found for the starting Pt(II) complex **1**, were observed and these are consistent with the observation of two signals in the  $^{195}\text{Pt}$  NMR spectrum of complex **3** at  $\delta -1797$  and  $-1806$  (two almost overlapping doublets each with  $^1J_{\text{PtP}} = 975$  Hz). These data would well establish the relative disposition of the different ligands on each Pt center as shown in Scheme 1 (with chirality at each Pt center), but it is not possible to use the present data to actually propose any “frozen” conformer(s) for complexes **3** and **4** resulting from rotation around one or two of the Pt–P bonds. As twice the “expected” number of signals was observed in the NMR spectra of complex **3**,

the formation of a statistical 1:1 mixture of two stereoisomers may be a reasonable explanation.

Complex **3** was also characterized by single-crystal X-ray diffraction analysis as illustrated in Fig. 4. It confirms the characterization of **3** as an octahedral diplatinum(IV) complex formed by oxidative addition of MeI to complex **1**. The coordination geometry around each platinum atom is octahedral with the Me group, the nitrogen atom and the *ortho* C of the ppy ligand, iodide and one P atom of the dppac. The three carbon atoms have *fac* geometry. The bond distances between platinum and carbon atom is  $2.089(14)$  Å for Pt1–C52. This value is considerably shorter than Pt1–C14 bond in the complex **2** (here the bond length is  $2.069(6)$  Å), indicating the conversion of Pt(II) to Pt(IV) product upon oxidative addition reaction. The iodide atom is located *trans* to the C atom of the ppy



**Fig. 5.** The reaction of complex **1** with excess MeI as monitored by  $^{31}\text{P}$  NMR spectroscopy in  $\text{CDCl}_3$  at room temperatures. Top: Pt(II) region. Bottom: Pt(IV) region. (a) pure **1**, (b) immediately after addition of excess MeI, (c) 8 min after addition of excess MeI, (d) 16 min after addition of excess MeI, (e) 36 min after addition of excess MeI, (f) 1 h after addition of excess MeI and (g) 100 min after addition of excess MeI. The peak assignments are shown.

ligand. The bond length of Pt1–C50 (*trans* to P1) and Pt1–C52 (*trans* to N1) are 2.102(6) and 2.089(14), respectively. The trend is in agreement with the observed  $^2J_{\text{PtH}}$  values in  $^1\text{H}$  NMR of complex **3**, showing lower *trans* influence of N atom as compared with that of P atom. The two methyl groups are *cis* disposed to one another (C50–Pt1–C52 92.0(4)°), indicating that the angles around the Pt center are rather close to the ideal angle of 90°. The aromatic N atom is coordinated to platinum center, which is in accord to the usual preference of the ppy ligand to form cyclometalated complexes, and is positioned *trans* to a Me group. The Pt1–P1 distance of 2.3886(14) Å in complex **3** is significantly longer than the Pt1–P1 distance of 2.2672(17) Å in complex **2**, consistent with the oxidation states of +4 and +2 for Pt atoms in complexes **3** and **2**, respectively. To the best of our knowledge this is the first X-ray structural determination of a diplatinum(IV) complex having dppac as bridging ligand.

### 3.3. Kinetic and mechanism of the reaction of complexes **1** and **2** with MeI

On the basis of the NMR and UV–vis spectroscopic studies, described below, a mechanism for reaction of complexes **1** and **2** with MeI is suggested as shown in Scheme 3.

#### 3.3.1. Monitoring the reaction of complex **1** with MeI by $^{31}\text{P}$ NMR spectroscopy

To investigate the details of mechanism of addition of MeI to diplatinum(II) complexes containing dppac as spacer ligand, the reaction of complex  $[\text{Pt}_2\text{Me}_2(\text{ppy})_2(\mu\text{-dppac})]$ , **1**, with excess MeI in  $\text{CDCl}_3$  was monitored by  $^{31}\text{P}$  NMR spectroscopy at room temperature. The spectra are shown in Fig. 5. Based on the results obtained from  $^{31}\text{P}$  NMR spectroscopy (and UV–vis spectroscopy, see next section), the complexes and intermediates suggested for the reaction sequence, described in Scheme 3, were assigned. The complexes **1** and **3** are characterized as described in the main text and Experimental section. Therefore the characteristic  $^{31}\text{P}$  NMR data for these complexes were used to indicate the complexes in the reaction sequence. Immediately after the addition of MeI at room temperature, apart from complex **1**, two Pt(II)–Pt(IV) species assigned as intermediates **A** and **B**, were detected. For intermediate **A**, two singlet signals, one at  $\delta$  14.0 (with  $^1J_{\text{PtP}} = 2047$  Hz close to that of the Pt(II)–Pt(II) starting complex **1**) and another one at  $\delta$  –24.0 with  $^1J_{\text{PtP}} = 639$  Hz, were observed. The latter value is typical for Pt(IV)–P coupling, but is higher than the corresponding value found for Pt(IV)–Pt(IV) complex **3** (with  $^1J_{\text{PtP}} = 628$  Hz); this

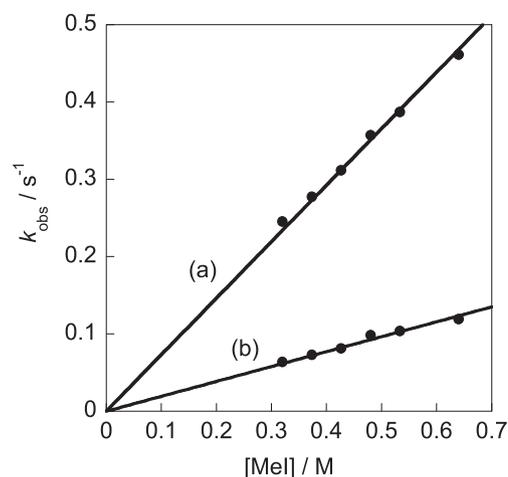


Fig. 7. Plots of first-order rate constants [(a) first step,  $k_{\text{obs}}/\text{s}^{-1}$ , and (b) second step,  $k'_{\text{obs}}/\text{s}^{-1}$ ] for the reaction of complex **1** with MeI at 30 °C versus  $[\text{MeI}]$  in  $\text{CHCl}_3$ .

is probably due to higher *trans* influence of Me as compared to that of the ppy-C ligand. For intermediate **B**, two singlet signals, one at  $\delta$  11.4 (with  $^1J_{\text{PtP}} = 1838$  Hz, located in the Pt(II) region) and another one at  $\delta$  –26.9, located in the Pt(IV) region with  $^1J_{\text{PtP}} = 988$  Hz, were observed. As the time was passing on, the signals due to the starting material **1**, and intermediates **A** and **B** were disappearing while those due to the intermediate **C** and final product **3** were growing. For the intermediate **C**, two signals in the Pt(IV) regions with expected  $^1J_{\text{PtP}}$  values (one signal at  $\delta$  –24.2, with  $^1J_{\text{PtP}} = 971$  Hz, another signal at  $\delta$  –24.3, with  $^1J_{\text{PtP}} = 980$  Hz) were observed. After 100 min, all the signals were completely disappeared and the final Pt(IV)–Pt(IV) product **7** was purely obtained.

#### 3.3.2. Kinetic studies using UV–vis spectroscopy

The kinetics of oxidative addition of MeI in  $\text{CHCl}_3$  to binuclear complexes **1** and **2** was studied by using UV–vis spectroscopy. In each case, an excess of MeI was used and disappearance of the MLCT band was followed to monitor the reaction. The change in the spectrum during a typical run is shown in Fig. 6. The absorbance–time curves monitored at 360 nm for the reaction of complex **1** (or 400 nm for complex **2**) with MeI in  $\text{CHCl}_3$  solution were obtained in the presence of large excesses of MeI, such that its concentration remained effectively constant in each experiment. The kinetic data were in accord with the occurrence of two sequential first-order

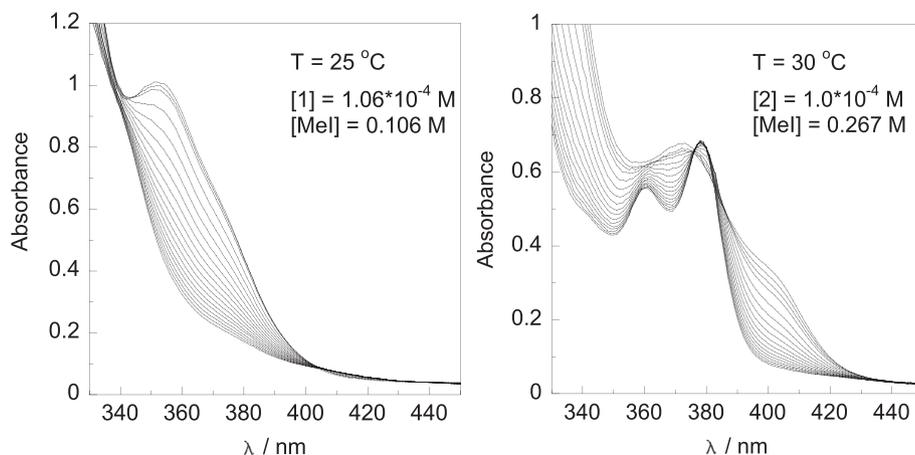


Fig. 6. Changes in the UV–vis spectrum during the reaction of complexes **1** (left) and **2** (right) with MeI in  $\text{CHCl}_3$ ; successive spectra were recorded at intervals of 30 s.

**Table 2**  
Rate constants<sup>a</sup> and activation parameters<sup>b</sup> for reaction of complexes **1** and **2** with MeI in CHCl<sub>3</sub>.

	Rate constants at different temperatures						$\Delta H^\ddagger/\text{kJ mol}^{-1}$	$\Delta S^\ddagger/\text{J K}^{-1} \text{mol}^{-1}$
	10 °C	20 °C	25 °C	30 °C	35 °C	40 °C		
<b>Complex 1</b>								
$10^2 k_2/(\text{L mol}^{-1} \text{s}^{-1})$	0.36	0.70	0.89	1.22	1.61	2.24	$41.9 \pm 0.9$	$-143 \pm 3$
$10^2 k'_2/(\text{L mol}^{-1} \text{s}^{-1})$	0.10	0.19	0.24	0.32	0.43	0.56	$38.9 \pm 0.8$	$-164 \pm 3$
<b>Complex 2</b>								
$10^2 k_2/(\text{L mol}^{-1} \text{s}^{-1})$	0.29	0.59	0.67	1.00	1.24	1.45	$37.4 \pm 2.2$	$-160 \pm 8$
$10^2 k'_2/(\text{L mol}^{-1} \text{s}^{-1})$	0.09	0.15	0.18	0.24	0.28	0.37	$31.1 \pm 1.1$	$-192 \pm 4$

<sup>a</sup> Estimated error in  $k$  values are  $\pm 5\%$ .

<sup>b</sup> Obtained from the Eyring equation.

reactions. Thus, the pseudo-first-order rate constants,  $k_{\text{obs}(1)}$  and  $k_{\text{obs}(2)}$ , for the two steps of the reaction were evaluated by nonlinear least-square fitting of the absorbance–time profiles to the biphasic first order equation (Eq. (1) and Fig. 1S).

$$\text{Abs}_t = \alpha \left[ \exp(-k_{\text{obs}(1)}t) \right] + \beta \left[ \exp(-k_{\text{obs}(2)}t) \right] + \text{Abs}_\infty \quad (1)$$

Both of the pseudo-first-order rate constants  $k_{\text{obs}(1)}$  and  $k_{\text{obs}(2)}$  show a variation with [MeI] that is linear with no intercepts (Fig. 7), showing a first-order dependence of the rate on the concentration of MeI in each step. Least-squares analysis gives the second-order rate constants ( $k_2$  for the first step and  $k'_2$  for the second step) and the results are given in Table 2. The findings are consistent with the two-step sequence indicated in the last section based on the <sup>31</sup>P NMR data (see Fig. 5) for the reaction of complex **1** with MeI. Therefore, each step of the reaction (see Scheme 3) obeys a simple second-order rate law (Eq. (2) for step 1 and Eq. (3) for step 2); note that  $k_2 > k'_2$  and the reproducibility of the data were remarkable ( $\pm 4\%$ ). The same method was used at other temperatures and activation parameters were obtained from the Eyring equation (see Eq. (4) and Fig. 2S). The data are shown in Table 2.

$$-d[\mathbf{1}]/dt = k_2[\mathbf{1}][\text{Me}] \quad (2)$$

$$-d[\mathbf{B}]/dt = k'_2[\mathbf{B}][\text{Me}] - k_2[\mathbf{1}][\text{Me}] \quad (3)$$

$$\ln\left(\frac{k}{T}\right) = \ln\left(\frac{k_B}{h}\right) + \frac{\Delta S^\ddagger}{R} - \frac{\Delta H^\ddagger}{RT} \quad k = k_2 \text{ or } k'_2 \quad (4)$$

The values of the entropy of activation were large and negative for both steps, which are in favor of a classical S<sub>N</sub>2 type mechanism. Based on these NMR and kinetic data, we suggest that as shown in Scheme 3, at the first step MeI attacks (via an electrophilic addition) on one of the electron rich Pt(II) centers of dimer complex **1** with an S<sub>N</sub>2 type mechanism with the rate constant  $k_2$  and the mix valence Pt(II)–Pt(IV) binuclear intermediate **A** is formed. The intermediate **A** would then quickly performs a facile *trans* to *cis* isomerization of the Pt(IV) center having Me and I in *trans* disposition to form the intermediate **B**. Intermediate **B** was then reacted with MeI in the second step again by an S<sub>N</sub>2 type mechanism, with the rate constant  $k'_2$  which is nearly 4 times smaller than the corresponding value for the rate constant of first step ( $k_2$ , see Table 2) to form intermediate **C** (Scheme 3). Then, a rapid *trans*–*cis* isomerization of the resulting Pt(IV) center was occurred to form the final product **3**.

As is clear from Table 2, in the two-step oxidative addition reactions studied in this work based on the qualitative <sup>31</sup>P NMR investigations and UV–vis calculation (see Supporting information for rate constants assignment using the Espenson's approach [68]), the step 2 appears to be slower than the step 1. In step 1, MeI attacks

the Pt(II) center of a Pt(II)–Pt(II) complex, **1**, while in step 2, MeI reacts with the Pt(II) center of a Pt(II)–Pt(IV) species (complex **B**). As such the reaction should be slower in the second step since in contrast to complex **1**, in species **2**, the adjacent Pt(IV) moiety reduces the nucleophilicity of the Pt(II) center through the bridging biphosphine group, dppac.

#### 3.4. Effect of CN ligand on the rate of oxidative addition reaction

As is clear from Table 2, the reaction rates of complex [Pt<sub>2</sub>Me<sub>2</sub>(bhq)<sub>2</sub>(μ-dppac)], **2**, with MeI in CHCl<sub>3</sub> at different temperatures for both steps, are at least 1.2–1.3 times slower than the corresponding reaction of complex [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)], **1**. For example, in step 1, MeI at 25 °C reacted nearly 1.3 times faster with [Pt<sub>2</sub>Me<sub>2</sub>(ppy)<sub>2</sub>(μ-dppac)] ( $k_2 = 5.36 \times 10^{-1} \text{ Lmol}^{-1} \text{ s}^{-1}$ ) than with **2** ( $k_2 = 4.01 \times 10^{-1} \text{ Lmol}^{-1} \text{ s}^{-1}$ ). This confirms the stronger donor ability of the ppy ligand as compared to that of the bhq ligand and shows that the platinum center in the ppy complex **1** is more electron rich than the platinum center in complex **2** toward electrophilic attacks. This is consistent with the trends found for <sup>1</sup>J<sub>PtP</sub> values in the <sup>31</sup>P NMR spectra of the analogous complexes **1** (2049 Hz) and **2** (2102 Hz). Therefore it may be concluded that the *trans* influence of metalated C atom of ppy ligand is greater than that of bhq ligand [57,69].

## 4. Conclusions

Two new binuclear complexes, [Pt<sub>2</sub>Me<sub>2</sub>(CN)<sub>2</sub>(μ-dppac)], in which CN = deprotonated 2-phenylpyridine (**1**) or benzo{h}quinoline (**2**) and dppac = bis(diphenylphosphino)acetylene, were synthesized. The related two-step reactions of complexes **1** and **2** with MeI were investigated by UV–vis spectrophotometry; the reaction of complex **1** with MeI was also monitored by <sup>31</sup>P NMR spectroscopy. On the basis of the results, both steps of MeI oxidative addition proceed by the classical S<sub>N</sub>2 type mechanism [70,71], as shown in Scheme 3, with large negative ΔS<sup>‡</sup> values. In the first step, the electron rich platinum center of the Pt(II)–Pt(II) starting complex **1** attacks the methyl carbon of MeI to form the intermediate **B** which is suggested to be a Pt(II)–Pt(IV) complex. The rate of the reaction of the platinum(II) center in complex **B** with MeI was considerably slower (by a factor of 4) than that for the platinum(II) center in the starting complex **1**, which is consistent with a significant electronic effect transmitted through the dppac bridging ligand. The Pt–P bond length for P atom *trans* to the C atom of ppy ligand is rather longer than that of the P atom *trans* to the C atom of bhq ligand. The difference, although modest, indicates that ppy ligand probably exerts a higher *trans* influence than the bhq ligand. This is in agreement with the values of <sup>1</sup>J<sub>PtP</sub> (observed in the <sup>31</sup>P

NMR spectra) and rates of MeI oxidative addition of the related complexes.

### Acknowledgments

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### Appendix A. Supplementary material

CCDC 928600 and 930544 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif).

### Appendix B. Supplementary data

Supplementary data related to this article can be found at <http://dx.doi.org/10.1016/j.jorganchem.2013.07.032>.

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