

Supplemental Materials: Intermolecular interactions in optical cavities: an ab initio QED study

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I. ADDITIONAL DATA FOR VAN DER WAALS BONDS

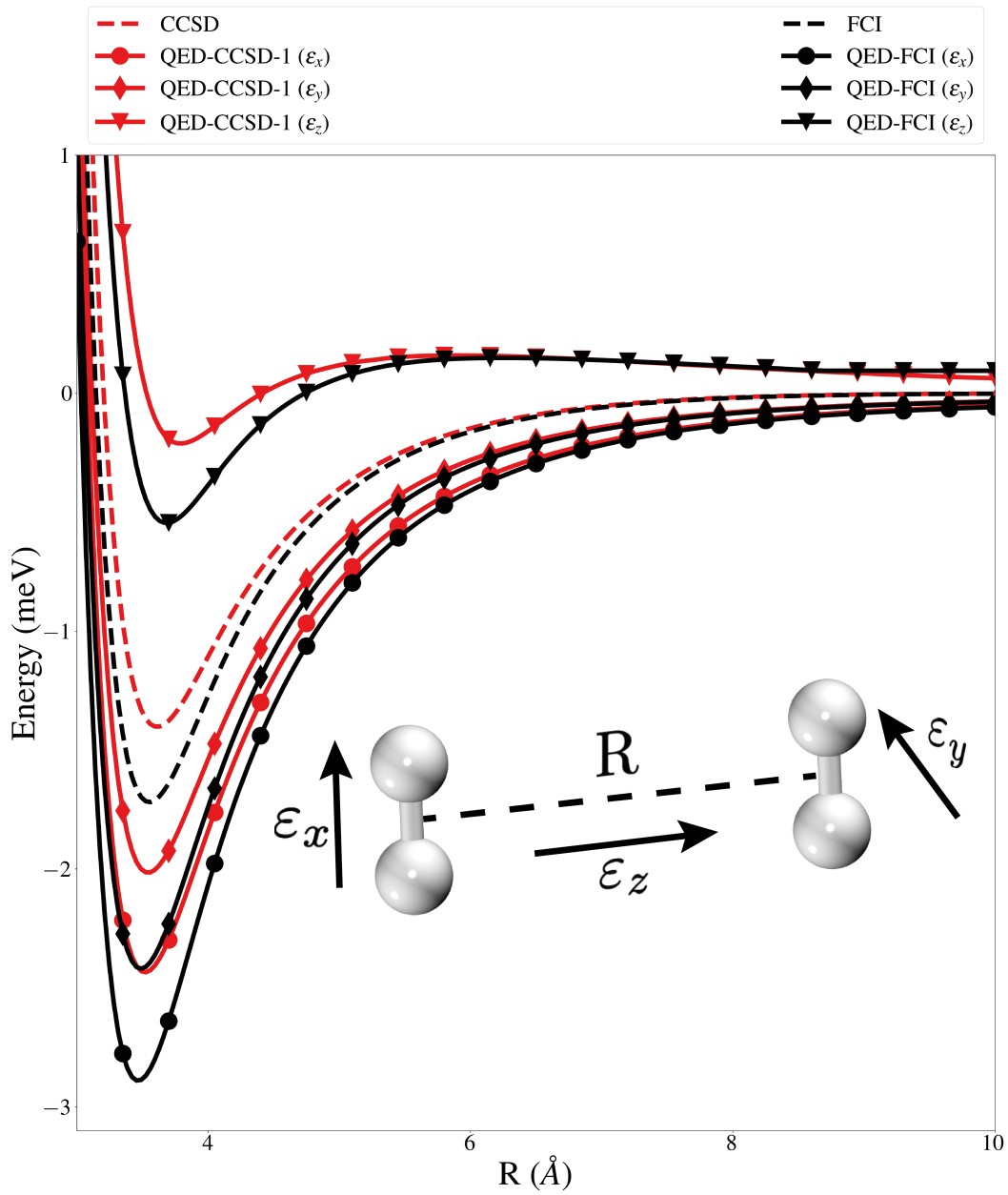


FIG. 1. $(\text{H}_2)_2$ polarization along H-H bond, orthogonal and along the chain.

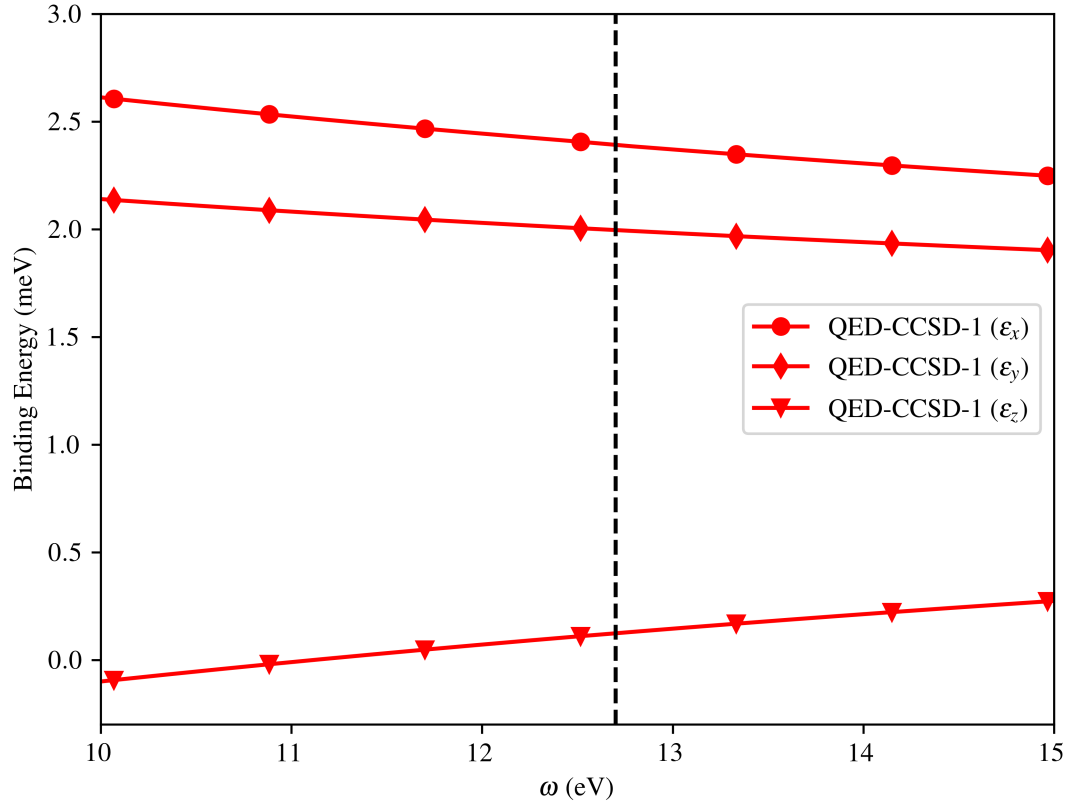


FIG. 2. Binding energy of two parallel H_2 molecules as a function of the photon frequency. Binding energies are calculated as energy differences between the value at 20 \AA and 3 \AA . The coupling value is set to $\lambda = 0.1$ and all three polarization directions have been presented. The black dashed line is the resonance frequency (12.7 eV).

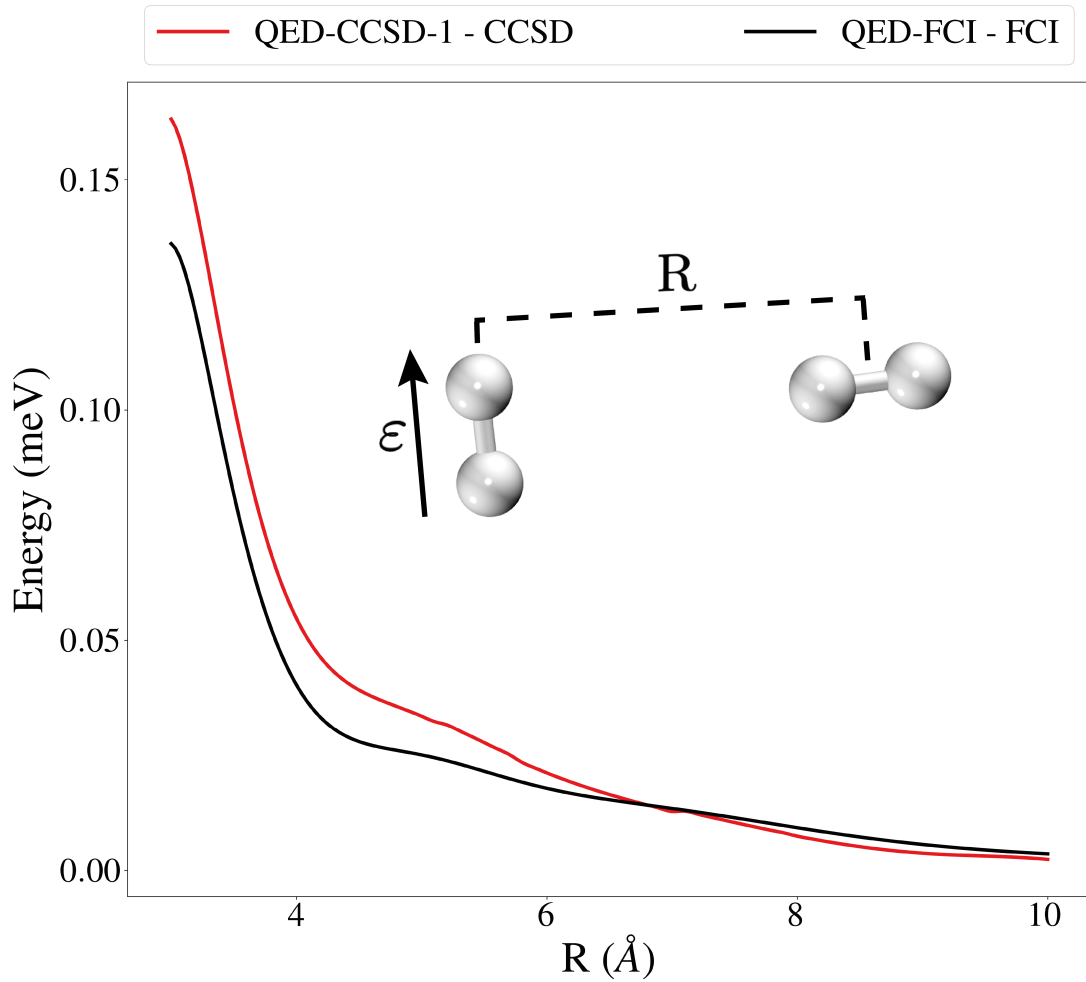


FIG. 3. Cavity induced effects on potential energy curves for two H₂ molecules in C_{2v} configuration stretching the distance between them. The polarization has equal components along the bond, chain and orthogonal directions with coupling $\lambda = 0.1$ and frequency $\omega = 12.7$ eV. The value at 200 Å has been subtracted.

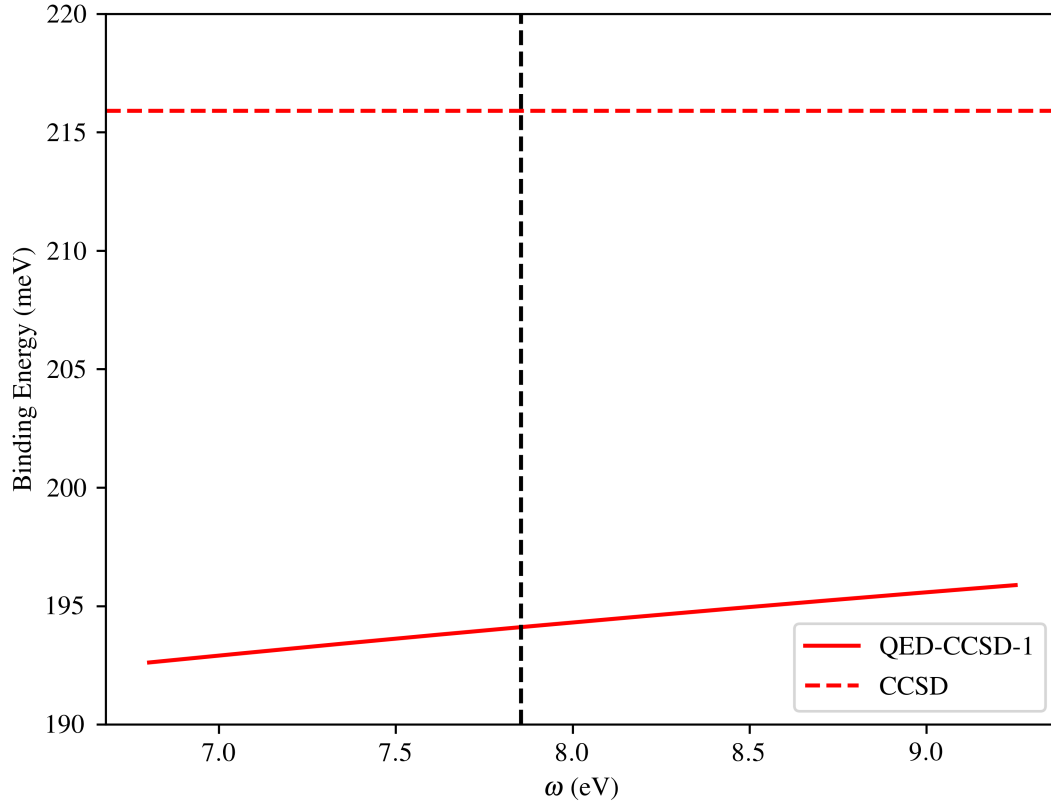


FIG. 4. Binding energy of the H_2O dimer as a function of the photon frequency. Binding energies are calculated as energy differences between the value at 20 \AA and 3 \AA distance between the oxygen atoms. The coupling value is set to $\lambda = 0.1$ and the polarization is parallel to the oxygens. The black dashed line is the resonance frequency (7.86 eV).