Tuning Single-Molecule Conductance in Metalloporphyrin-Based Wires via Supramolecular Interactions

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In memory of Professors Nongjian Tao and Fausto Sanz

Abstract: Nature has developed supramolecular constructs to deliver outstanding charge-transport capabilities using metalloporphyrin-based supramolecular arrays. Herein we incorporate simple, naturally inspired supramolecular interactions via the axial complexation of metalloporphyrins into the formation of a single-molecule wire in a nanoscale gap. Small structural changes in the axial coordinating linkers result in dramatic changes in the transport properties of the metalloporphyrin-based wire. The increased flexibility of a pyridine-4-yl-methanethiol ligand due to an extra methyl group, as compared to a more rigid 4-pyridinemethyl linker, allows the pyridine-4-yl-methanethiol ligand to adopt an unexpected highly conductive stacked structure between the two junction electrodes and the metalloporphyrin ring. DFT calculations reveal a molecular junction structure composed of a shifted stack of the two pyridinic linkers and the metalloporphyrin ring. In contrast, the more rigid 4-mercaptopyridine ligand presents a more classical lifted octahedral coordination of the metalloporphyrin metal center, leading to a longer electron pathway of lower conductance. This work opens to supramolecular electronics, a concept already exploited in natural organisms.

Introduction

The concept of Supramolecular Electronics arises as a result of blending the studies of organic crystalline systems and the field of conducting polymers.[1] In parallel, in the last decade, new crystal structure information on natural biomolecular wires has fascinated the scientific community by revealing the way nature exploits supramolecular electronics using arrays of axially coordinated metalloporphyrins to create highly efficient molecular conduits.[2,3] Supramolecular electronics also provides a unique opportunity to study the mechanobiology of electrical signaling, as another key aspect of the emerging field of mechanochemistry[4] in biological systems.[5] Metalloporphyrins have been extensively studied as molecular wires owing to a number of appealing properties such as high chemical stability and conjugation, modular metal center and rich supramolecular chemistry.[6–10] Metalloporphyrins have been chemically connected to metal electrodes either by directly lying flat on the metal surfaces via π-orbital interactions between the metal and the porphyrin ring,[11,12] or by covalent electrode/molecule attachment through porphyrin ring substituents.[13–18] Although the latter results in a robust, straightforward method to wire oligo-
porphyrins between two electrodes, the inclusion of such anchoring Scheme precludes the exploitation of other sources of supramolecular interactions that might lead to the formation of more efficient electron pathways already exploited in the natural biomolecular homologous wire. We have recently reported a novel way to form highly conductive metalloporphyrin wires by coordinating axial positions of the metalloporphyrin ring allowing to orient the ring plane perpendicularly to the electron pathway (main junction axis).\[19–21\] This is possible thanks to the highly axial coordinative affinity of both metal center and porphyrin ring to strong Lewis bases. Such axial ligands act as anchoring groups or linkers,\[22\] mimicking the common natural schemes exploited in the chemistry of photosynthetic and transmembrane electron transport.\[2,3,23\]

In this contribution, we aim to rationalize the conductance landscape of a metalloporphyrin-based supramolecular wire under mechanical stress by systematically introducing structural changes of both the axial coordinative ligands and the porphyrin chemical substitution. To this goal, we built single-molecule junctions using an STM-break junction approach of Co\(^{3+}\)-porphyrins (Figure 1) with different phenyl substitutions; unsubstituted (P), 5,15-diphenyl (DPP) and 5,15-dibisphenyl (DBP) substituted metalloporphyrins, employing thiol-functionalized electrodes with two different pyridine compounds as axial coordinative linkers; a pyridine-4-yl-methanethiol (PyrMT)\[19–21\] and a 4-pyridinethiol (PyrT). We show that this slight structural change in the axial ligands (differing by one methyl group) results in pronounced changes in the dominant supramolecular interactions that lead to the final molecular wire geometry, and ultimately dictates its final transport properties. We perform extensive DFT structural and charge transport simulations of the Co-porphyrin supramolecular wire using the two axial ligands to help visualizing the most plausible junction configurations in each of the studied cases.

**Results and Discussion**

**Single-molecule charge transport of the Co-DPP/PyrMT and Co-DPP/PyrT assemblies**

The STM-BJ technique\[24\] was used to form and measure the conductance of individual Co-DPP dissolved in a non-polar organic medium when they get trapped between the two Au electrodes of a STM junction functionalized with either PyrMT or PyrT linkers. The electrode functionalization is done ex-situ by exposing the electrodes to a solution of each thiolated linker (see details in Supplementary Information (SI) section 1 and 6). We employed the dynamic STM-BJ method, referred as tapping,\[24\] where the STM tip electrode is driven repeatedly in and out of contact with the substrate electrode (see details in SI section 6). During the retraction cycle, individual Co-DPP dissolved in the working media can spontaneously span the electrodes gap forming a transient molecular junction. Characteristic plateaus-like features show up in the current versus separation (retraction) curves as a result of molecular wire formation and breakdown (see representative ones in Figure 2 insets). Typically, several hundred (up to a thousand) of those retraction curves displaying plateaus features are selected and accumulated in 1D and 2D semi-log conductance histograms (details in SI section 2 and 6) resulting in prominent peaks, which provide the most probable single-molecule conductance values. Figure 2a and b show the corresponding 1D histograms for Co-DPP junctions employing the PyrMT and PyrT linkers, respectively (see corresponding 2D histograms in SI section 2). Both junctions present multiple conductance features (peaks I to III) that are attributed to stable pyridine/metalloporphyrin interactions as the STM gap expands giving rise to different supramolecular associations. Experiments in the absence of either Co-DPP (PyrMT and PyrT only) or linkers (Co-DPP only) show no evident peak features within the same conductance range (in SI section 4.2). The DPP (metal-free metalloporphyrin) junctions in the presence of both PyrMT and PyrT linkers show both the absence of the highest peak I conductance signature (see SI Section 4.1), which reveals that the peak I feature results from the axial coordination to the metal center in both cases, while peaks II and III are the result of pyridine/porphyrin ring interactions.\[19\] The final junction’s geometries are, however, unknown. The multiple conductance features observed in all the above studied cases bring several findings: (i) the more flexible PyrMT linker allows a larger number of possible stable junction’s geometries, (ii) the geometries achieved with the flexible PyrMT linker are two orders of magnitude more conductive than those with the more rigid PyrT, and (iii) two well differentiated interactions, via metal or via porphyrin ring, take place with both linkers.
Co-DPP/PyrMT and 2020 G linker, on the other hand, computes a much lower X\text{G} shows a much larger propensity to “lie down” on PyrT and 3.50 \times 10^{-5} G_0, which is three orders of magnitude below the experimental value (2.82 \times 10^{-5} G_0). Such discrepancy is too large to be associated with DFT approximations, including PBE + U functional corrections for the well-known underestimated highest (lowest) occupied (unoccupied) molecular orbitals (HOMO–LUMO) energy gap in the GGA functionals, which usually results in even larger (overestimated) conductance values. We then performed a detailed structural analysis of the most likely geometries of the linker/Co-DPP/linker adduct in a constrained tunneling gap. The geometry of the PyrMT and PyrT linkers alone on the electrode surface was optimized using a many-body approach to include the dispersion term.[26,27] The PyrMT shows a much larger propensity to “lie down” on the electrode surface (Figure 3a), scoring 14.0 kcal mol^{-1} more stable than a “standing up or tilted” geometry. The PyrT linker, on the other hand, computes a much lower 6.0 kcal mol^{-1} difference between both conformations, adopting a slightly higher tilt angle (Figure 3b), and suggesting higher likelihood of finding the linker in a “lifted” geometry when forming part of a compact monolayer, as the ones prepared in the experiments (see SI section 6). Moreover, a standing up geometry of an adsorbed PyrT in a compact monolayer on gold has been previously suggested from STM imaging.[28,29] The above computational results (Figure 3a,b) suggest the PyrMT and PyrT might interact with the Co-DPP in a π-π stacking and in a more axial coordination, respectively. To corroborate this scenario, we have also performed ellipsometry measurements on a molecular layer of the form linker/Co-DPP/linker for both PyrMT and PyrT linkers on a clean Au surface (Figure 3c). The resulting thickness values are consistent with the formation of a Au/PyrMT/lying down)/Co-DPP/PyrMT, scoring the lowest measured thickness (11.6 Å), and a thicker (13.0 Å) Au/PyrT(lifted)/Co-DPP/PyrT, thus supporting our initial hypothesis: in the presence of Co-DPP, the adsorbed PyrMT and PyrT linkers on the Au surface adopt a lying down and lifted geometries, respectively. Note that the ellipsometry analysis corresponds to the structure of the linker monolayer only (Figure 3c) and that the final geometry upon molecular junction formation (two electrodes) might differ from the one in the ellipsometry experiment (one electrode). Notwithstanding, the ellipsometry data highlights a completely distinct pyridine interaction to the metal electrodes, resulting in the more rigid PyrT linker significantly decoupled from the metal surface (thicker measured layer) in comparison with the more flexible PyrMT (thinner measured layer), despite the smaller molecular length of the latter.

DFT-optimized structures for the whole linker/Co-DPP/ linker supramolecular junction for the two linkers are shown in Figure 3d and e, and the computed conductance values from the corresponding zero-energy transmission functions are 6.84 \times 10^{-5} G_0 and 2.46 \times 10^{-4} G_0 for the PyrMT and PyrT junctions, respectively, in excellent agreement with the experimental data corresponding to the peak I features, 2.82 \times 10^{-5} G_0 and 3.50 \times 10^{-4} G_0, respectively (Figure 2).
major supramolecular interactions lead in each case to a completely different junction; the PyrMT linker forms a π-stacked conformation coordinating the metal center, while the PyrT preferably coordinates the metal center in a quasi-fully standing up fashion. Figure 3d and e configurations are then ascribed to the most likely molecular wire configurations leading to transmissions given by the peaks I in Figure 2a and b. Note that the experimental (dynamic) tilted angle of the PyrT might differ from the suggested quasi-fully standing up conformation in Figure 3e, as deduced from the ellipsometry data. With independence of the final exact tilted angle in the molecular junction, a major impact on conductance due to the PyrT-pyridine decoupling from the Au surface is observed. We conclude here that the larger flexibility of the PyrMT promotes the stabilization of the final junction structure through π-π stacking interactions, while the rigid PyrT tends to adopt a more orthogonal coordinative geometry. The shorter conduction path in the PyrMT junction justifies its larger transmission. It is also important to remark the unusual coordination geometry of the metalloporphyrin with the PyrMT ligand, which evidences the interplay between linker-linker neighbor interactions and linker-Au surface interactions (Figure 3a), the latter stabilizing the observed final “lying down” conformation for the PyrMT ligand.
Now we turn our attention to the conductance features II and III in Figures 2a and b, originating from interactions between the linkers and the porphyrin ring, not involving the metal center. To this aim, we first perform additional experiments using an unsubstituted porphyrin (Co-P) and its metal-free (P) homologous (Figure 4). The Co-P junction using PyrMT linkers (Figure 4a) show two distinguishable molecular conductance signatures I and II at $9.31 \times 10^{-3} G_0$ and $3.01 \times 10^{-4} G_0$ respectively. Same experiment with P (Figure 4b) yields a unique low conductance feature II centered at $2.24 \times 10^{-4} G_0$ close to the equivalent feature II in Figure 4a, which tentatively leads us to same previous peak assignment: π-stacked pyridine/metal for peak I (equivalent to signature I in Figure 2) and pyridine/porphyrin ring for peak II interactions (equivalent to either II or III in Figure 2). The overall conductance decreases for all homologous signatures in the Co-P junction as compare to the Co-DPP (see summarizing Table 1), being especially acute for conductance feature II ($\approx 30 \times$). This conductance changes are due to the different phenyl substitution in both Co-DPP and Co-P cases, and they are tentatively ascribed to two different electrical contributions: (i) the phenyl substitution might bring the energy of the HOMO frontier orbital closer to the Fermi level (see the projected density of states (PDOS) in SI section 5) via electron donation by resonance and reduce the energy barrier for the transmitted electrons assuming off-resonance tunneling through the HOMO level (see also phenyl contribution to the HOMO in Figure S5.8). Such resonance effect is expected to be more pronounced for the pyridine/porphyrin ring interaction (feature II), as experimentally observed. And (ii), the close electrode/phenyl proximity in the highly constrained Co-DPP junctions (Figure 3d) could also increase the electrode/porphyrin contact area (coupling) giving rise to an increased conductance.

The absence of one of the low conductance signatures (II, III) in Figure 4a when compared to the Co-DPP results with the same linker (Figure 2a) evidences also the active role of the phenyl porphyrin substitutions in the formation of the supramolecular junction. This pyridine/phenyl interaction is also confirmed by the same measurements performed on junctions based on a 5,15-dibisphenylporphyrin (DBPP), with bi-phenyl substitutions in the porphyrin ring (see SI section 4.3). The DBPP junctions with PyrMT linkers display additional conductance signatures (up to 4 overlapping peaks are visible) as compare to the two observed in DPP (peaks II

![Figure 4](image-url)
and III in SI section 4.1), which demonstrates the additional accessible interaction sites brought by each phenyl ring.

The Co-P junctions employing a PyrT linker (Figure 4c) shows a unique conductance feature whose value is within the same order of magnitude as the conductance signature I in the Co-DPP junctions with the same linker (Figure 2b), and therefore, ascribed to an axial “lifted” coordination of the pyridine linker to the metal center. The small discrepancy (≈2.3 ×, see Table 1) in both Co-P and Co-DPP cases is again tentatively ascribed to the electronic donating resonant effect of the phenyl substituents in the porphyrin ring. The absence of conductance signatures II and/or III using PyrT linkers for Co-P (Figures 4c,d) demonstrate that the enhanced flexibility of the PyrMT linker, which presents a conductance signature for the junction with the unsubstituted free-metal porphyrin P (Figure 4b), readily interact with the porphyrin ring thanks to its more accessible π-π stacking orientation (Figure 3a). When both the metal center and the phenyl substitutions are removed, the PyrT is unable to establish any stable interaction with the porphyrin backbone P resulting in a silent conductance histogram (Figure 4d).

**Computational analysis of the low conductance features II and III**

We again computed DFT-optimized junction geometries to visualize the most plausible interactions that lead to the observed II and III conductance features for the PyrMT and PyrT linkers, where the metal is partially or not participating in the junction formation. Figure 5 summarizes the minimal-energy DFT junction configurations whose computed conductance values lie within the range of the experimental ones. Starting with the PyrMT, Figures 5a and b show two stable configurations corresponding to the replacement of one and two PyrMt-porphyrin ring interactions, respectively, by PyrMT-phenyl interactions with the Co-DPP. These interactions arise from effective π-π stacking between the pyridine moiety of the PyrMT linker and the phenyl substitutions of the porphyrin. The calculated conductance values for these two optimized geometries are $6.81 \times 10^{-3}$ and $4.50 \times 10^{-3}$ G, in agreement with the experimental values of $9.75 \times 10^{-3}$ and $3.02 \times 10^{-3}$ G (Table 1), and they are also consistent with a dynamic picture of consecutive formation of more extended conformations as the junction is elongated during the tip retraction (SI Figure S3.1), where Figure 5a,b conformations are able to span slightly larger electrode-electrode gap.

![Figure 5](image_url)

**Table 1**: Summary of most probable conductance values (expressed as $10^{-4}$ G) of all observed signatures for each porphyrin/linker combinations.

<table>
<thead>
<tr>
<th>Linkers</th>
<th>Porphyrins</th>
<th>Co-DPP</th>
<th>DPP</th>
<th>Co-P</th>
<th>P</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Peaks</td>
<td>I</td>
<td>II</td>
<td>III</td>
<td>I</td>
</tr>
<tr>
<td>PyrMT</td>
<td>282</td>
<td>97.5</td>
<td>30.2</td>
<td>–</td>
<td>80.8</td>
</tr>
<tr>
<td>PyrT</td>
<td>3.5</td>
<td>–</td>
<td>0.17</td>
<td>–</td>
<td>–</td>
</tr>
</tbody>
</table>

![Figure 5](image_url)
separations (see SI section 3 for a comparison of the experimental gap separations and the DFT geometries distances of the assemblies). In the absence of phenyl substitutions (Co-P and P), we found effective supramolecular interactions between the pyridine moiety of the two PyrMT and the pyrrolic ring of the porphyrin (Figure 5c) characterized by the conductance feature II in Figure 4a,b. Note that this conformation does not imply interactions with the metal center and that they also seem to be hindered in the Co-DPP case by the presence of the phenyl substitutions, where the pyridine-phenyl interaction dominates. The computed conductance in Figure 5c configuration yields $5.5 \times 10^{-4} \, G_0$, in good agreement with the experimental value for the Co-P/PyrMT signature II (Table 1). As for the PyrT, the optimized geometries for the Co-DPP/PyrT system suggest two plausible scenarios where either one or two metal/pyridine coordination(s) is(are) replaced by \( \pi-\pi \) phenyl-pyridine interactions which are enabled thanks to the dihedral rotation of the phenyl substituent (Figure 5d). The calculated conductance values are $4.21 \times 10^{-5}$ and $1.37 \times 10^{-5} \, G_0$, respectively, both close to the corresponding experimental III signature, $1.71 \times 10^{-5} \, G_0$ (Table 1). The total DFT energies for both configurations indicate the interaction with two phenyl substituents (Figure 5d right panel) is $12.6 \, \text{kcal mol}^{-1}$ more stable suggesting this one as the most plausible scenario for conductance feature III in Figure 2b. This assumption is also supported by the results corresponding to the free-metal DPP, where the homologous configuration with two phenyl/pyridine interactions (Figure S5.7) is also $17.5 \, \text{kcal mol}^{-1}$ more stable, leading to a calculated conductance of $1.48 \times 10^{-5} \, G_0$, close to the experimental value of $1.56 \times 10^{-5} \, G_0$ (Figure S4.1b). The absence of the peak II signature and of any conductance signature in Figure 4c and d, respectively is then explained by the inability of the PyrT linker to establish effective \( \pi-\pi \) interactions with the pyrrolic ring.

The summarizing Figure 6 maps out all the supramolecular landscape leading to the proposed configurations in our metalloporphyrin-based single molecule junctions, pinpointing each of the computed supramolecular geometries to every observed single molecule conductance feature. The generally found good agreement between computed and experimental conductance values reinforces the adequacy of DFT + U corrected methods in the studies of supramolecular electronics. Figure 6 picture conceptually opens to new ways of designing nanoscale molecule wires exploiting well-known supramolecular interactions, paving the way to Supramolecular Electronics. We also expect this work to serve as a platform to study charge transport in biological moieties system exploiting very similar supramolecular interactions to produce scenarios for long-range electron transfer.

**Figure 6.** Schematic diagram of the supramolecular landscape for all studied molecular junctions. The conductance values are represented on the x-axis in $G_0$ units (solid lines are the experimental values, stripped lines the theoretical values) for both linkers (PyrMT and PyrT) and the four studied porphyrin systems (Co-DPP, Co-P, DPP, and P). Schematic representations of the structural models confirmed by DFT are drawn for each conductance signature. The color legend represents equivalent interactions across all junctions: blue corresponds to the both pyridine linkers interacting with the metal center, green is one pyridine interacting with the metal center and the other with a side phenyl group, red is both pyridine interacting with side phenyl rings, and yellow corresponds to two pyridine linkers interacting with the porphyrin pyrrolic ring.
Conclusion

Concluding, we have studied the formation of single-molecule electrical contacts in a tunneling junction exploiting the rich axial coordination landscape in metalloporphyrin systems using pyridine-based linkers. We demonstrate that changes in the linker flexibility result in strikingly different supramolecular interactions between the pyridinic linker and the pyrrolic porphyrin ring, leading to radically different molecular junction geometries. As summarized in Figure 6, an extra methyl group in a PyrMT linker, as compared to a PyrT linker, confers extra conformational degrees of freedom to the pyridine group resulting in the formation of pyridine/porphyrin π-π stacking conformations, as opposed to the more orthogonal axial coordination geometries occurring with the more rigid PyrT linker. The supramolecular wires conductance resulting from these two distinct geometries differ by two orders of magnitude, being the π-π stacking conformations more conductive. As the molecular junction is mechanically stretched, more extended supramolecular configurations are identified where either the pyrrolic ring or/and the phenyl side groups of the metalloporphyrin readily provide with additional interacting points to the pyridine linkers, allowing switching to new pathways where the metal center is partially or not participating in the final supramolecular junction structure.

These results demonstrate the large conductance tunability of a molecular wire via tweaking its internal supramolecular interactions and present a novel platform to investigate the fascinating, yet unknown, field of the mechanism of electron transport in complex biomolecular structures. This porphyrin-based supramolecular wire also brings a new synthetic testbench to study the molecular mechanisms supporting the exceedingly long-range (from mesoscopic to micrometer range) charge transport in bacteria nanowires exploiting similar supramolecular interactions.[55]

Acknowledgements

This research was supported by the Spanish Ministerio de Ciencia, Innovación y Universidades (grants PGC2018-093863-B-C21, MAT2016-77852-C2–1-R and MDM-2017-0767) and the ERC Grants Fields4CAT (grant 772391) and Tmol4TRANS (grant 724981). I. D.-P. thanks King’s College London for start-up funds support. A.M.R. thanks the Ministerio for a predoctoral FPI grant. E.R. thanks Generalitat de Catalunya for an ICREA Academia award and for the SGR2017-1289 grant. N.A.-A. and A.G.-C. thank the Generalitat de Catalunya for the grant 2017SGR1277 and the Severo Ochoa Program for Centers of Excellence in R&D (SEV-2015-0496). A.G.C. thanks Ministerio de ciencia e Innovacion for Ramon y Cajal grant (RYC-2017-22910). The authors acknowledge the general facilities of the University of Barcelona (CIT-UB) and the computer resources, technical expertise and assistance provided by the Barcelona Supercomputing Centre.

Conflict of interest

The authors declare no conflict of interest.

Keywords: biomolecular electronics - density functional calculations - metalloporphyrins - single-molecule junctions - supramolecular electronics


Manuscript received: May 19, 2020
Accepted manuscript online: July 24, 2020
Version of record online: August 24, 2020